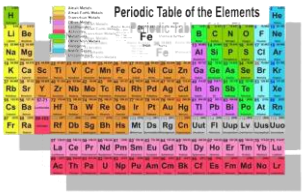
Name:_____ **Date:**_____ **Worksheet**
Chemistry



Periodic Trends

DESCRIBE the following terms:

electronegativity -

ionization energy -

atomic radius -

DIRECTIONS: Choose the atom in each of the following groups with the **GREATER** electronegativity.

_____ C, O _____ As, Bi _____ Na, K _____ F, Br _____ I, O

DIRECTIONS: Arrange the following elements according to INCREASING electronegativity.

a) C, N, O, F, Ne b) S, Se, Cl, Ar, Sr

DIRECTIONS: Arrange the following elements according to INCREASING first ionization energies.

a) Sr, Be, Mg, Ca, Kr b) F, Te, I, Xe, S

DIRECTIONS: Tell which ion or atom has the LARGER atomic radius in each of the following sets.

_____ a) F^-, Ne, Na^+ _____ b) O^+, O, O^- _____ c) Mg^+, Sr^+, Ca^+

_____ d) Ca, As, Cr _____ e) F, Kr _____ f) Sr^{2+}, Rb^+, Kr

DIRECTIONS: Draw the Lewis dot symbols for the following atoms.

iodine potassium boron cobalt silicon

"To improve is to change; to be perfect is to change often."
- Winston Churchill

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