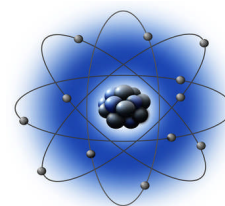


Name: _____ Date: _____
Chemistry

Review



Atomic Models and Structure

DESCRIBE the following:

atomic number -

ion -

isotope -

cation -

anion -

mass number -

FILL IN THE BLANKS with the best term.

_____ in any sample of a compound the masses of the elements are in the same proportions

_____ when two elements combine to form more than one compound, the fixed amounts of one element will combine with the other in a ratio of small whole numbers

_____ if a system is left to itself will move toward disorder

LIST the six main atomic models.

_____	_____
_____	_____
_____	_____

LIST the three major subatomic particles and the charge.

	particle name	charge (relative)	mass (relative)
1.	_____	_____	_____
2.	_____	_____	_____
3.	_____	_____	_____

WRITE the correct person for the following statements.

- _____ proposed the name for the electron
- _____ actually performed the Gold-foil experiment
- _____ named the proton and developed the Nuclear atomic model
- _____ proved the existence of the electron
- _____ developed the Solid Sphere atomic model
- _____ determined the charge of the electron
- _____ developed the Plum-pudding atomic model
- _____ developed the concept of the atomic number
- _____ proved the existence of the neutron

FILL IN the chart below.

Substance	Symbol	Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
Nitrogen		7			7	
Potassium		19	39			
	Cu		64			

WRITE the nuclide symbol for the following:

_____ Si that has 15 neutrons in the nucleus

_____ Pt that has a mass number of 190

LIST the number of protons, neutrons, electrons for the following isotopes.

isotopes	${}^3\text{He}^{2+}$	${}^{207}\text{Pb}$	${}^{36}\text{Cl}^-$	${}^{11}\text{B}^{3+}$
protons				
neutrons				
electrons				

WRITE the nuclide symbol for the following.

_____ **24** protons; **27** neutrons; **21** electrons

_____ **33** protons; **42** neutrons; **36** electrons

“Success is in our daily habits.”